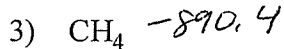
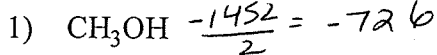


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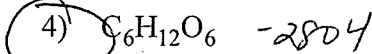
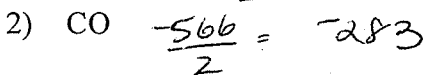
Key

Table I Practice Problems

4 1) Based on the *Heats of Reaction at 101.3 kPa and 298 K* chemistry reference table, the most energy would be released when oxygen reacts completely with 1.0 mole of

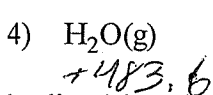
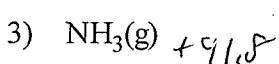
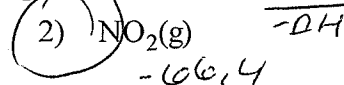
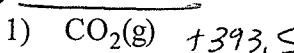


2

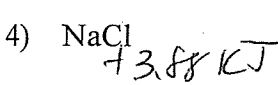
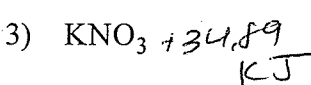
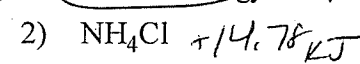
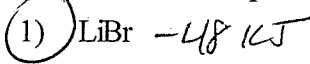


2 2) According to the *Heats of Reaction at 101.3 kPa and 298 K* chemistry reference table, the decomposition of which compound would be an exothermic reaction?

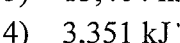
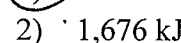
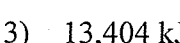
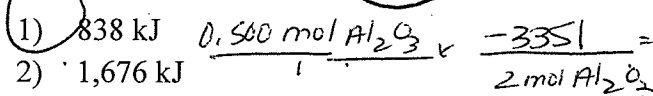
switch signs



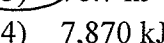
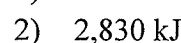
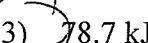
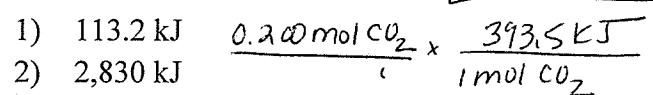
1 3) According to the *Heats of Reaction at 101.3 kPa and 298 K* chemistry reference table, the dissolving of which salt is accompanied by the release of energy?



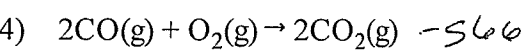
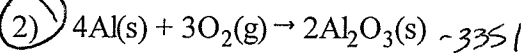
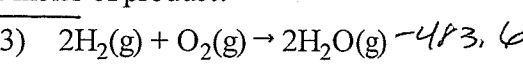
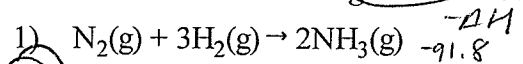
1 4) Based on the *Heat of Reaction at 101.3 kPa and 298 K* chemistry reference table, how many kilojoules of heat are given off when 0.500 mole of $\text{Al}_2\text{O}_3(\text{s})$ is formed from its elements?



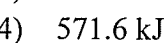
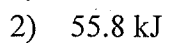
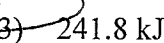
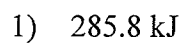
3 5) Based on the *Heats of Reaction at 101.3 kPa and 298 K* chemistry reference table, how many kilojoules of heat are given off when 0.200 mole of $\text{CO}_2(\text{g})$ is formed from its elements?



2 6) Which reaction releases the greatest amount of energy per 2 moles of product?



3 7) According to the *Heats of Reaction at 101.3 kPa and 298 K* chemistry reference table, what is the total amount of energy released in the formation of one mole of $\text{H}_2\text{O}(\text{g})$?



$\frac{483.6}{2} =$