

Naming Practice 2

1. calcium carbonate CaCO_3 $\text{Ca}^{2+} \text{CO}_3^{2-}$
2. sodium hydride NaH $\text{Na}^{+1} \text{H}^{-1}$
3. beryllium phosphate $\text{Be}_3(\text{PO}_4)_2$ $\text{Be}^{2+} \text{PO}_4^{3-}$
4. iron(III)hydroxide $\text{Fe}(\text{OH})_3$ $\text{Fe}^{3+} \text{OH}^{-1}$
5. zinc chloride ZnCl_2 $\text{Zn}^{2+} \text{Cl}^{-1}$
6. magnesium acetate $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$ $\text{Mg}^{2+} \text{C}_2\text{H}_3\text{O}_2^{-1}$
7. silver nitrate AgNO_3 $\text{Ag}^{+1} \text{NO}_3^{-1}$
8. copper(II)nitrate $\text{Cu}(\text{NO}_3)_2$ $\text{Cu}^{2+} \text{NO}_3^{-1}$
9. aluminum sulfide Al_2S_3 $\text{Al}^{3+} \text{S}^{2-}$
10. ammonium hydroxide NH_4OH $\text{NH}_4^{+1} \text{OH}^{-1}$
11. manganese(IV)oxide MnO_2 $\text{Mn}^{+4} \text{O}^{2-}$
12. calcium chloride CaCl_2 $\text{Ca}^{2+} \text{Cl}^{-1}$
13. lead(II)chlorate $\text{Pb}(\text{ClO}_3)_2$ $\text{Pb}^{2+} \text{ClO}_3^{-1}$
14. mercury(I)oxalate $\text{Hg}_2\text{C}_2\text{O}_4$ $\text{Hg}_2^{2+} \text{C}_2\text{O}_4^{2-}$
15. potassium phosphide K_3P $\text{K}^{+1} \text{P}^{3-}$
16. barium nitride Ba_3N_2 $\text{Ba}^{2+} \text{N}^{3-}$
17. chromium(III)carbonate $\text{Cr}_2(\text{CO}_3)_3$ $\text{Cr}^{3+} \text{CO}_3^{2-}$
18. gold(I)sulfite Au_2SO_3 $\text{Au}^{+1} \text{SO}_3^{2-}$
19. lithium bromide LiBr $\text{Li}^{+1} \text{Br}^{-1}$
20. calcium carbide Ca_2C $\text{Ca}^{2+} \text{C}^{4-}$